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Language: English. Information about the book "P.C. Rakshit, Physical Chemistry" can be viewed on the publisher's web site. P.C. Rakshit, Physical Chemistry Introduction This book has three objectives. Firstly, to acquaint the student with the basic concepts of thermodynamics. Secondly, to help in understanding the concept of entropy and give him/her a lucid understanding of the third law of thermodynamics. Thirdly, to improve the basic understanding of thermodynamics by introduction of a number of other concepts which have been covered up to date. The book also has appendices, I, II, and III, with the help of which the reader will be able to understand the topics much more clearly and to make much better sense out of the concepts dealt with. The course is suitable for the following readers: post graduate and post graduate students and teachers in this field. Contents Section 1.1 General Remarks: Introduction, Concepts, Sections, Chap.1. Principle of Entropy and Entropy as a Concept of Thermodynamics. Chap.2. Entropy for Thermodynamics. Entropy for Ideal gas. State functions. State function of a system. Principle of entropy conservation. Entropy function. Functions as function of states and entropy. Definition of molar entropy. T.F.Dentropy function and the third law. Equilibrium, the first law and the second law. H.P.atkins and Equilibrium. General laws of thermodynamics. Conservation of energy. Heat. Section 1.2 Introduction: Introduction to the topic. Section 1.3 Conservation of energy: The first law of Thermodynamics. Section 1.4 Equilibrium: Introduction, Equilibrium, Equilibrium, G.M.Watson. General definition of equilibrium. Simple definition of equilibrium. Principles of Gibbs. Van der Waals and Gibbs. Section 1.5 The second law of thermodynamics: Introduction, The second law. The second law of thermodynamics. Entropy, entropy generation, entropy as a physical property. Equilibrium and the second law. The connection between the third law and the second law. The third law. Randomness and entropy. Section 1.6 Free energy: Introduction, Free energy. Definition of free energy. The change of free energy of a system: the relation between entropy and free energy. Entropy change f3e1b5768c

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